

Name: _____
Period: _____ Date: _____

Ion Worksheet

1. The atomic number of oxygen is 8 and the mass number is 16.

A. Draw an oxygen atom.

B. What is the charge of the oxygen atom you just drew? _____

C. The oxygen atom can gain two electrons (to make 8) and become an oxygen ion. Draw the ion.

D. What is the charge of the oxygen ion?(Hint: It is not neutral) _____

E. What is the symbol for the oxygen ion?

2. The atomic number of lithium is 3 and the mass number is 6.

A. Draw the lithium atom.

B. What is the charge of the lithium atom you just drew? _____

C. The lithium atom loses one electrons and become lithium ion. Draw the ion.

D. What is the charge of the lithium ion? (Hint: It is not neutral) _____

E. What is the symbol for the lithium ion?

3. The atomic number of nitrogen is 7 and the mass number is 14. Draw the N^{-3} ion.

4. Fill in the blanks. Following are ions.

Element	Atom No.	# of p+	# of e-	Charge	Symbol
Sodium	11	11	10	+1	Na^{+1}
Chlorine		17		-1	
Sulfur	16				
Calcium		20		+2	
Iodine	53				I^{-1}
zinc		30	28		
Oxygen					
Magnesium					

5. Write the formula for the ionic molecule made when these ions combine.

	Cl^{-1}	O^{-2}	S^{-2}	N^{-3}	Br^{-1}
Na^{+1}					
Ag^{+1}					
K^{+1}					
Zn^{+2}					
Al^{+3}					
Mg^{+2}					
Fe^{+3}					
Sn^{+4}					

6. Name these compounds.

- a. KCl _____
- b. PbI _____
- c. MgO _____
- d. AgCl _____
- e. $ZnCl_2$ _____
- f. $AlBr_3$ _____
- g. Fe_2O_3 _____
- h. $CoCl_2$ _____
- i. KBr _____
- j. $BaCl_2$ _____
- k. MgS _____